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Fridrich, Bálint; Stuart, Marc C.A.; Barta, Katalin

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Selective Coupling of Bioderived Aliphatic Alcohols with Acetone Using Hydrotalcite Derived Mg–Al Porous Metal Oxide and Raney Nickel

Bálint Fridrich,[†] Marc C. A. Stuart,[‡] and Katalin Barta^{*,†}

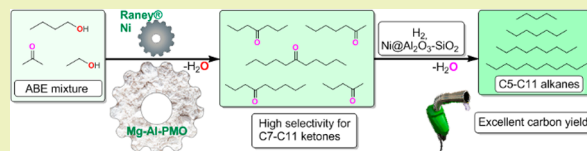
[†]Stratingh Institute for Chemistry, University of Groningen, Nijenborgh 4, 9747 AG Groningen, The Netherlands

[‡]Department of Electron Microscopy, Groningen Biomolecular Sciences and Biotechnology Institute, University of Groningen, Nijenborgh 7, 9747 AG Groningen, The Netherlands

S Supporting Information

ABSTRACT: Fermentation of sugars to the so-called ABE mixture delivers a three component mixture of shorter chain oxygenates: acetone, *n*-butanol and ethanol. In order to convert these into liquid transportation fuels that are analogous to the currently used fossil energy carriers, novel catalytic chain elongation methods involving C–C bond formation are desired. Herein we report on a simple, non-noble-metal-based method for the highly selective coupling of 1-butanol and acetone into high molecular weight (C7–C11) ketones, as well as ABE mixtures into (C5–C11) ketones using the solid base Mg–Al–PMO in combination with small amount of Raney nickel. Upon hydrodeoxygenation, these ketones are converted to fuel range alkanes with excellent carbon utilization (up to 89%) using Earth abundant metal containing catalysis.

KEYWORDS: Bioethanol, 1-Butanol, ABE mixture, Raney nickel, Porous metal oxides



INTRODUCTION

Lignocellulosic biomass is a globally available, carbon-neutral, renewable resource that does not compete with the food supply and that could, in principle, provide suitable and sustainable alternatives to the fossil-based chemicals currently made.^{1–3} It could also be used to produce liquid transportation fuels that are, in their chemical structure, analogous to the generally utilized petroleum derived sources and are therefore fully compatible with existing infrastructures.⁴ However, the conversion of naturally abundant carbohydrates to fuel-range hydrocarbons (>C8) is not straightforward, because the central monosaccharide building block, D-glucose only consists of six carbon atoms.⁵ Thus, beside appropriate downstream processing of the sugar platform, also novel methods for C–C bond formation and subsequent deoxygenation are desired.^{6–8}

Fermentation of sugars to produce ethanol is a historically well-established method^{9,10} that is also significant today for the production of first and second generation bioethanol.^{11,12} Another interesting biochemical transformation of cellulosic or noncellulosic feedstocks is the so-called ABE (acetone, butanol, ethanol) fermentation,^{13,14} which provides a three component mixture of these small molecular weight oxygenates.¹⁵ Pioneering work from the Toste group has presented an integrated approach that combines fermentation of sugars to ABE mixture that is subsequently catalytically converted to yield C5–C11 ketones and ultimately high energy density (30.2–35 MJ/L), fuel range alkanes with excellent carbon balance.¹⁶ Their approach recognized the potential of acetone to readily undergo aldol-condensation with the aldehydes in situ generated from the primary alcohols in the ABE mixture. The required dehydrogen-

ation, aldol condensation and hydrogenation sequence was performed using the physical mixture of Pd@C and K₃PO₄ to yield chain elongated ketones. Next, the group introduced recyclable hydrotalcite-based catalyst doped with various metals.¹⁷ Later, a synergistic effect was found in a Pd/Cu doped hydrotalcite that allowed for high product selectivity at 200 °C in the gas phase.¹⁸ An in depth study of the hydrogenation and decarbonylation of butanol over various PdCu bimetallic catalysts comprising different supports was conducted.¹⁹ Further, it was observed that HTC doped with up to 2.5% Cu (prepared by impregnation), upon calcination, leads to excellent reaction rates.²⁰ Other groups also found excellent results in this and related reactions, such as the palladium catalyzed chain elongation of diluted aqueous ABE mixture,²¹ or the related iPrOH, EtOH, BuOH mixture.²² The copper catalyzed alkylation of acetone with BuOH²³ and the catalytic dehydration of ABE mixture over γ -alumina catalyst show the versatility of this chemical platform.²⁴ For instance, Lee and co-workers have used Pd@C with CaO as a solid base for the conversion of ABE mixture in a batch reactor at 180 °C without any added solvent to obtain a mixture of ketones and their corresponding alcohols with 78% yield based on acetone.²⁵ The mixture was further converted one-pot to the corresponding alkane mixture using 34.5 bar H₂ at 270 °C, with 74.7% overall yield based on acetone. Onyestyák and co-workers designed a flow reactor containing two separated beds for the consecutive

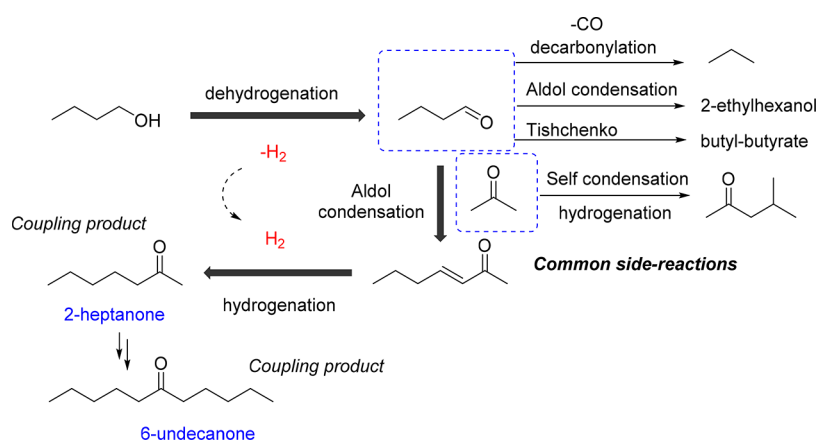
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Scheme 1. Key Reaction Steps and Main Side Reactions in the Catalytic Conversion of 1-Butanol and Acetone to C7 and C11 Ketones



coupling and hydrodeoxygenation sequence to produce C5–C7 alkanes from ethanol-acetone mixture.²⁶ Tan, Shen and co-workers have introduced a Ni-MgO-SiO₂ catalyst with improved water resistance for the highly selective conversion of solventless ABE mixtures to C5–C15 ketones and alcohols at 240 °C.²⁷

Herein we describe a simple catalyst system consisting of Mg–Al–PMO and Raney nickel, for the very efficient coupling of acetone and 1-butanol²⁸ as well as the selective coupling of ABE mixture to long chain ketones at 200 °C and hydrodeoxygenation of these over non-noble-metal catalysts to the corresponding fuel range alkanes with excellent carbon balance.

EXPERIMENTAL PROCEDURE

Catalyst Preparation. The HTC (hydrotalcite) catalyst precursor was prepared by a coprecipitation method, according to literature procedures.^{29–31} A solution containing MgCl₂·6H₂O (0.15 mol, 30.05g), AlCl₃·6H₂O (0.05 mol, 12.07g) in deionized water (0.2 L) was slowly added to an aqueous solution (0.2 L) of Na₂CO₃ (0.05 mol, 5.3 g) at 60 °C under vigorous stirring. The pH was carefully maintained between 9 and 10 by adjusting with frequent additions of an aqueous solution of NaOH (1 M). The mixture was vigorously stirred for 72 h at 60 °C. After cooling to room temperature, the suspension was filtered, and the solid was washed with deionized water and resuspended into a solution of Na₂CO₃ (2 M), which was stirred for 24 h at 40 °C. After, the catalyst precursor was filtered and washed with deionized water until the washings were chloride-free (tested with AgNO₃ solution). The solid was dried at 120 °C overnight, and the hydrotalcite precursor (HTC) was obtained as a white powder (14.7 g). The corresponding porous metal oxide (PMO) was obtained after calcining a portion of the HTC material at 460 °C for 24 h. Elemental analyses were performed on a PerkinElmer instrument (Optima 7000DV) after full solubilization of the PMO catalysts in diluted nitric acid.

The various doped hydrotalcites (Cu20PMO, Cu10Ni10PMO, Cu10Zn10PMO, Ni20PMO, Co20PMO) were prepared by replacing a portion of the Mg²⁺ ions with additional metal dopants, otherwise following the procedure above. In Cu20PMO, Ni20PMO and Co20PMO, 20 mol % of the Mg²⁺ were replaced by the corresponding two valent metal ions Cu²⁺, Ni²⁺ and Co²⁺ respectively. In the Cu10Ni10PMO and Cu10Zn10PMO, 10 mol % of Mg²⁺ was replaced by Cu ions whereas 10 mol % was replaced by Ni and Zn, respectively. To demonstrate the successful incorporation of the two-valent transition metal ions, the composition of PMO catalysts was investigated by ICP analysis, which typically resulted in an elemental composition that closely matched the theoretical one (for a detailed description, see Supporting Information, Section 4.1 Table S15). In order to independently confirm the formation of hydrotalcite structures, powder XRD spectra for all HTCs and for the corresponding PMOs were recorded (see Supporting Information, Section 4.2 for details). Indeed,

this analysis confirmed the formation of the envisioned hydrotalcite structures in all samples.

Representative Procedure of a Catalytic Coupling of Butanol with Acetone. First, 15 mg of Raney nickel slurry was placed in a Swagelok stainless steel microreactor (10 mL) equipped with a Teflon coated magnetic stirring bar followed by a rapid addition of 1-butanol (5.46 mmol, 0.5 mL), acetone (2.7 mmol, 0.2 mL), and dodecane (0.22 mmol, 50 μL, internal standard). Next, 200 mg of Mg–Al–PMO and heptane (2×1 mL) solvent was added. The reactor was sealed and placed in an aluminum block preheated at the desired temperature. After the indicated reaction time, the microreactor was cooled down with an ice–water bath and subsequently carefully opened. The liquid sample as well as the catalysts were quantitatively transferred to a 15 mL centrifuge tube. Additional MeOH (2×1 mL) was used to wash the reactor and recover all catalyst residues. After centrifugation, the solution was transferred into a glass vial, washed four times with 2 mL MeOH and analyzed by GC-FID (Hewlett-Packard 6890 series equipped with a HP-5 capillary column) in order to determine the 1-butanol and acetone conversion and selectivity of the products. Compounds were also identified by GC–MS and the injection of pure reference standards for the comparison of retention times in the GC column.

For ABE mixture, acetone (2.9 mmol, 0.21 mL), butanol (4.6 mmol, 0.42 mL), and ethanol (1.2 mmol, 70 μL) were used as substrate; otherwise, the procedure was identical. The ratio of A:B:E is kept at 2.3:3.7:1 molar ratio as expected from the mixture accessible directly from ABE fermentation.

Catalytic Hydrodeoxygenation of Long Chain Ketones to Alkanes. The mixture of chain elongated ketones obtained following the previous procedure, (or model compounds) in hexane (~10 mL) were loaded into a 15 mL autoclave equipped with a magnetic stirring bar and 100 mg Ni@Al₂O₃–SiO₂ catalyst. The reactor was additionally purged with H₂ three times then pressurized to 40 bar, the autoclave was placed in a metal block preheated at 220 °C and the reaction proceeded for 12 h. After, the reactor was cooled down by compressed air and ice–water bath, respectively. The content of the reactor was transferred to a 15 mL centrifuge tube and the reactor was washed with additional amounts (5× 0.5 mL) of cold hexane. The catalyst was separated by centrifugation, a sample from the liquid phase was injected to GC–MS and parallel to GC-FID and the carbon yield was calculated based on carbon number relatively to the intensity of the internal standard. More details on product analysis are given in the Supporting Information, Section 1.2 and 1.5.

RESULTS AND DISCUSSION

Inspired by the above-described prior art and following our interest in the use of hydrotalcite derived doped porous metal oxides for the conversion of renewables involving hydrogenation³² or hydrogen transfer processes,^{29,33} as well as diverse C–C bond formation reactions including the Guerbet coupling

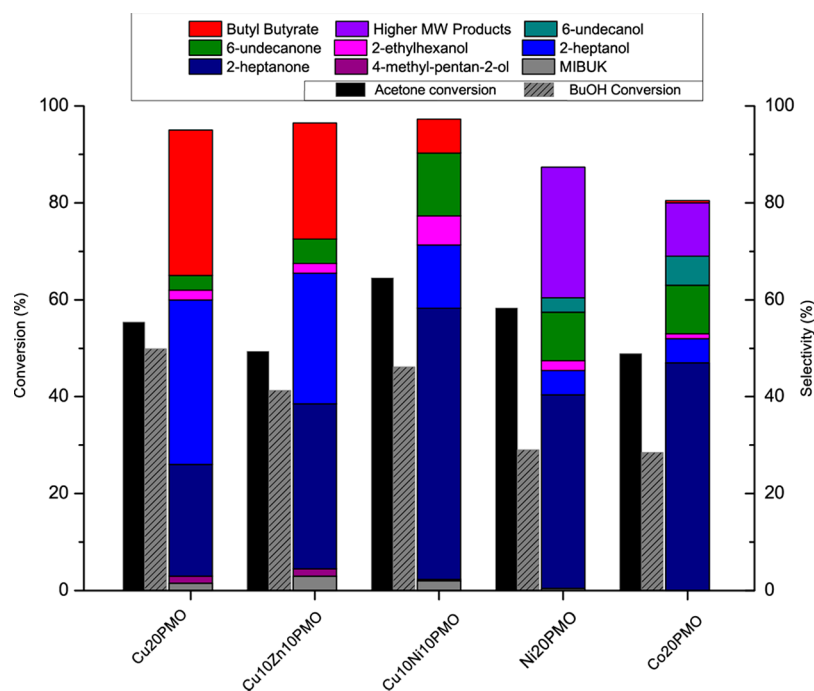


Figure 1. Screening of various PMO catalysts containing different metal dopants in the catalytic coupling of acetone with 1-butanol. Right: Acetone and 1-butanol conversion. Left: Product selectivity by GC-FID. Reaction conditions: 200 mg catalyst, acetone (2.70 mmol, 0.2 mL), 1-BuOH (5.46 mmol, 0.5 mL), acetone–butanol molar ratio 1:2, heptane (2.1 mL, solvent), dodecane (0.22 mmol, 50 μ L, internal standard), 20 h, 200 $^{\circ}$ C. For all data in table format, see the Supporting Information, Table S1.

of ethanol to butanol³⁴ and the coupling of challenging alcohol mixtures with cyclopentanone,³⁵ we set to design a simple, entirely noble-metal-free catalyst system for the transformation of ABE mixture to long chain ketones and subsequently to fuel-range alkanes. We have first focused our attention on finding an appropriate catalyst for the coupling of 1-butanol and acetone, two main components of the ABE mixture. As shown in Scheme 1, the transformation of acetone and 1-butanol to long chain ketones requires efficient dehydrogenation of the primary alcohol, followed by a cross aldol condensation with acetone and subsequent hydrogenation of the double bond in the corresponding aldol condensation product. A number of side reactions, previously also identified by Toste¹⁸ and others, exist in this system, making the selective formation of 6-undecanone challenging.

First, we have selected a number of hydrotalcite derived solid bases, high surface area porous metal oxides (PMO) comprising various dopants (Cu, Ni, Zn, Co), as starting point for our investigation (Figure 1, Table S1). We expected that the nature of the dopant will influence reactivity since in our earlier study on the Guerbet reaction we observed significant differences in the extent of dehydrogenation, aldol condensation as well as the Tishchenko pathway, depending on the PMO composition. Gratifyingly, there was considerable conversion of both acetone and 1-butanol with all PMO catalyst tested, and a significant amount of the one side coupling product 2-heptanone together with the corresponding alcohol 2-heptanol. Indeed, with diverse metal dopants, the extent of side reactions appeared different. As expected from earlier studies, Cu20PMO exhibited significant amounts of butyl-butyrate and a similar behavior was observed with Cu10Zn10PMO (Table S1, entries 1 and 2). A related behavior was observed by Goulas et. al who reported a significant esterification side reaction with HTC supported Cu catalyst prepared by impregnation.²⁰ Interestingly, the Cu10Ni10PMO

containing the same amount of Cu dopant, but Ni instead of Zn resulted in much less butyl butyrate (7%) and higher amount (56%) of 2-heptanone as well as (13%) long chain ketone 6-undecanone (Table S1, entry 3). The PMO containing Ni or Co both resulted in very similar substrate conversion values, up to ~60% ketones and more significant amount of higher Mw products (Table S1, entries 4 and 5). Although among these catalysts the best results were achieved with the Cu10Ni10PMO and the conversion values and low selectivity to the double coupling products were not yet satisfactory.

A very high selectivity to (C7–C11) ketones was described by Goulas et. al reporting the beneficial use of a CuPd bimetallic catalyst prepared by incipient wetness impregnation on a commercially available HTC catalysts.¹⁸ The group has identified decarbonylation as one of the major source of side reactions, catalyzed by the Pd contained in the catalysts. Because the Cu containing catalysts were giving significant Tishchenko activity, and none of the doped PMO tested gave sufficient conversion levels, we turned our attention to the use of simple Mg–Al–PMO obtained directly by calcination of a typical hydrotalcite that is known to efficiently catalyze aldol condensation processes.³⁶ Furthermore, instead of Pd we have selected Raney nickel that is a highly efficient transfer hydrogenation catalyst, but is expected to have less decarbonylation activity. Very recently, Zhang and co-workers have shown that cyclopentanol can be converted to C10 and C15 oxygenates in the presence of Raney Ni and a PMO with a different Mg/Al ratio (2:1 compared to 3:1 in our case).³⁷

Combination of Raney Nickel and PMO for Increased Activity. Gratifyingly, when we have attempted the coupling reaction between acetone and 1-butanol in the presence of 15 mg Raney Ni and 200 mg Mg–Al–PMO, excellent results were obtained (Figure 2, Table S2, Entry 1). Almost full substrate conversion was seen alongside high selectivity (61.5%) of 6-

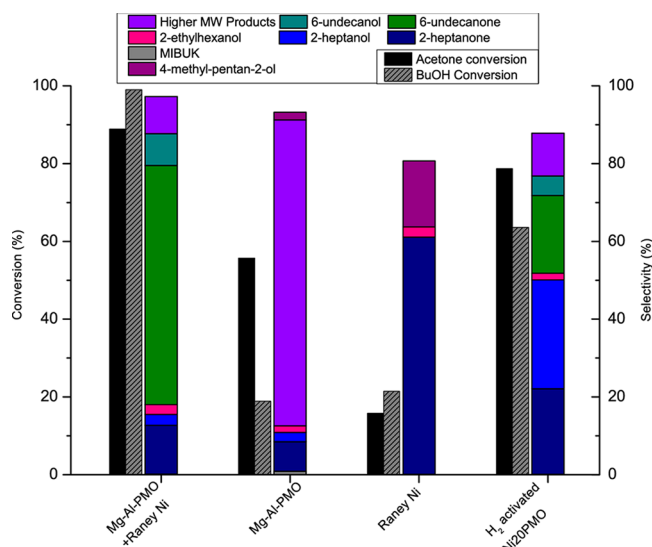


Figure 2. Evaluating various Ni containing catalyst in the catalytic coupling of acetone with 1-butanol. Left: Acetone and 1-butanol conversion. Right: Product selectivity by GC-FID. Reaction conditions: 200 mg catalyst, acetone (2.70 mmol, 0.2 mL), 1-BuOH (5.46 mmol, 0.5 mL), acetone–butanol molar ratio 1:2, heptane (2.1 mL, solvent), dodecane (0.22 mmol, 50 μ L, internal standard), 20 h, 200 $^{\circ}$ C. For all data in table format, see the Supporting Information, Table S2.

undecanone and the combined selectivity for C7 and C11 ketones and alcohols as desired chain elongated products was over 86%. In terms of side reactions, a small amount of higher molecular weight side products were detected, but no butylbutyrate was formed. The superior activity of this catalyst can be explained by the fact that Raney Ni is a more efficient dehydrogenation and hydrogenation catalyst, compared to the various metal dopants in the PMO previously tested, present in their oxide form. When the Mg–Al–PMO was used alone, the color of the reaction mixture turned orange and the formation of higher M_w products, originating from acetone self-condensation were largely predominant, as expected. On the other hand, when only Raney Ni was used, low substrate conversion was seen and a small quantity of 2-heptanone was found. Interestingly, also 4,4-dibutoxybutane was detectable that likely originates from the acetalization of the butanal intermediate formed upon dehydrogenation, a pathway apparently more dominant in the absence of aldol condensation, supporting the major role of Raney Ni in catalyzing dehydrogenation. This mixture of catalysts was investigated by TEM measurement, combined with EDX. As expected, the phase boundaries of the physical mixture were clearly observable even after reaction (for description, see the Supporting Information, Section 4.3). Interestingly, when the Ni20PMO was first preactivated at 500 $^{\circ}$ C in hydrogen atmosphere and was evaluated in the reaction after preactivation, the amount of 6-undecanone increased. This hints at the fact, that Ni(0) together with a basic porous metal oxide, are essential for good activity. Related Ni–Al hydrotalcite derived catalysts are known for CO₂ methanation and Ni–Mg–Al hydrotalcite derived catalysts for ammonia decomposition.^{38–41} For all these systems, the formation of different-sized Ni nanoparticles was observed upon activation in H₂ stream. It was previously reported that Ni nanoparticles, which were well isolated and immobilized in the porous metal oxide matrix displayed higher thermostability, while also being easily accessible to the reactants.⁴⁰

In our hands, the freshly prepared, not activated Ni20PMO revealed homogeneously distributed nickel and a porous structure as evidenced by TEM and EDX analysis (Figure S13). Upon activation at 500 $^{\circ}$ C in H₂ flow, formation of particles was visible by TEM analysis; however, their sizes were too small (1–4 nm) to discriminate them from the matrix by elemental mapping (Figure S14). Therefore, from the same batch of porous metal oxide precursor, we prepared a new batch of activated catalyst at 600 $^{\circ}$ C in H₂ flow. Indeed, in this case the formation of larger, Ni-rich particles (7–11 nm) was observed (Figures S15 and S19).

The catalyst, which was activated at 600 $^{\circ}$ C, showed slightly better conversion and similarly good selectivity toward the chain elongated products (Table S2, entry 5).

Next, both Ni-PMO (activated at 500 and 600 $^{\circ}$ C, respectively) were analyzed by TEM and EDX after reaction (Figure S18 vs S19).

The flexible incorporation of Ni into the PMO structure synthetically, will open possibilities for potential future fine-tuning of catalyst composition and preparation methods, because the amount of Ni in the PMO during coprecipitation method can be carefully adjusted.

Effect of Various Reaction Parameters. Next, we conducted a series of experiments by gradually increasing the substrate amounts by maintaining the total volume of 2.8 mL of the reaction mixture (Figure 3, Table S3). Experiments were

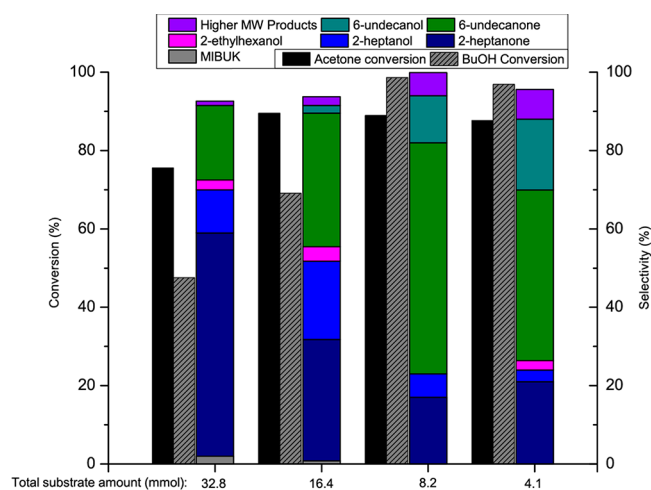


Figure 3. Evaluating the effect of increasing substrate loading in the catalytic coupling of acetone with 1-butanol. Left: Acetone and 1-butanol conversion. Right: Product selectivity by GC-FID. Reaction conditions: 200 mg Mg–Al–PMO and 15 mg Raney Ni catalyst, acetone–butanol molar ratio 1:2, dodecane (0.22 mmol, 50 μ L, internal standard), 20 h, 200 $^{\circ}$ C. For all data in table format, see the Supporting Information, Table S3.

conducted between 4.1 and 32.8 mmol total substrate loading. Interestingly, though the acetone conversion did not significantly change, the 1-butanol conversion and 6-undecanone selectivity gradually decreased with an increasing amount of substrate loading. Gratifyingly, the neat conditions and relatively high substrate loading also resulted in an excellent selectivity for C7–C11 ketones and alcohols (combined 86.3%), albeit with lower conversion values.

In the next experiments, the acetone to butanol ratio was varied as 1.9, 2.0, 2.1 (Figure S1 and Table S4). The 1:2 acetone to butanol ratio resulted in the highest amount of double

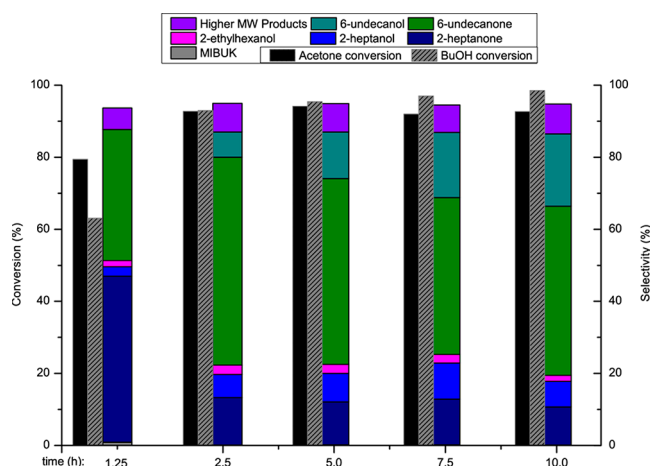


Figure 4. Product formation profiles in the catalytic coupling of acetone with 1-butanol for 10 h. Left: Acetone and 1-butanol conversion. Right: Product selectivity by GC-FID. Reaction conditions: 200 mg Mg–Al–PMO and 15 mg Raney Ni catalysts, acetone (2.70 mmol, 0.2 mL), 1-BuOH (5.46 mmol, 0.5 mL), acetone–butanol molar ratio 1:2, hexane (2.1 mL, solvent), dodecane (0.22 mmol, 50 μ L, internal standard), 200 $^{\circ}$ C. For all data in table format, see the Supporting Information, Table S7.

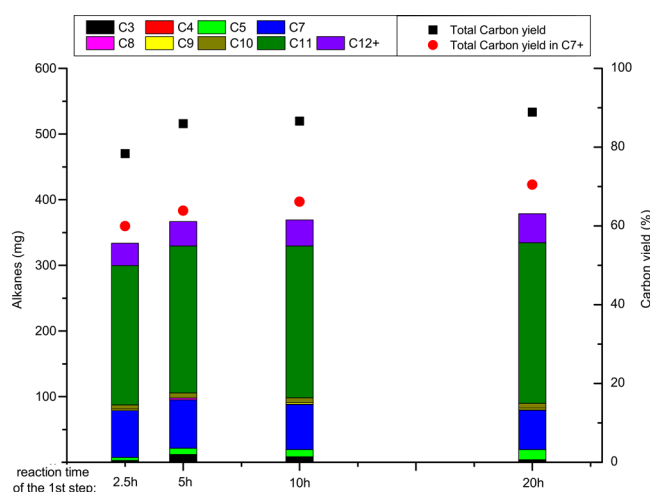


Figure 5. Results of the catalytic HDO of product mixtures obtained upon catalytic coupling of acetone with 1-butanol for different reaction times. Left: Weight of alkanes obtained based on chain length. Right: Carbon yields. Reaction conditions: 100 mg catalyst: Ni@Al₂O₃–SiO₂, dodecane (0.22 mmol, 50 μ L, internal standard), 220 $^{\circ}$ C. Values determined by FID, using internal standard. For all data in table format, see the Supporting Information, Table 1 and Table S12.

coupling step (see the Supporting Information, Section 2.1.2) and second the leftover acetone in the product mixture prior to HDO was transformed to volatile propane (Table S10, Figure 5).

The Ni/SiO₂–Al₂O₃ catalyst was also used in our earlier work for the HDO of lignocellulose derived product mixtures at 250 $^{\circ}$ C. It displayed excellent robustness and recyclability without coking in the HDO of model compound 4-propylcyclohexanone for 20 consecutive cycles, converting 12.4 g (88.5 mmol) substrate with just 200 mg catalyst.³⁵ Therefore, we postulate that this catalyst will be highly suitable for a continuous flow setup involving HDO of products originating from the coupling of ABE mixtures.

Coupling of ABE Mixture and Hydrodeoxygenation.

Under the conditions found ideal for the AB mixture, we have explored the catalytic coupling of the ABE mixture at 200 $^{\circ}$ C (Figure 6, Table S8). The molar ratio of the substrates was kept (2.3:3.7:1), mimicking the real fermentation mixture. Gratifyingly, full 1-butanol and ethanol conversion and 87.3% acetone conversion was seen after 20 h (581 μ mol·g cat^{−1}·h^{−1}) and an excellent 88.4% selectivity toward oxygenates was achieved. These oxygenates comprised ketones (77.4%, out of which 57.6% were C7+ ketones) and a smaller amount of alcohols (11% out of which 7.5% were C7+ alcohols). Moreover, the extent of side reactions was minimal, since only 9.6% of high molecular weight side products and 0.2% MIBUK were obtained.

Next, catalytic HDO of this mixture was performed and the carbon yields of alkanes (Figure S8.) were calculated based on an internal standard. According to this, an excellent 84% total carbon yield was achieved (Table 2, Table S13). The carbon yield of C7+ alkanes was found to be slightly lower compared to the value seen with the AB mixture, because acetone mono or dialkylation with EtOH yields C5–C7 ketones respectively, which do not contribute to the C7+ region after HDO.

CONCLUSIONS

Here we describe that the solid base Mg–Al–PMO in combination with small amount of Raney Ni provides high selectivity in the coupling of 1-butanol with acetone as well as ABE mixture to longer chain ketones. The extent of common side reactions such as ester formation and decarbonylation could be minimized. Notably, this catalyst system provides high selectivity to chain elongated ketones (73%) and alcohols (12%) as well as a total C7+ ketones (62%) and alcohols (9.1%) in the coupling of acetone and 1-butanol at 200 $^{\circ}$ C and it functions comparably well at 180 $^{\circ}$ C. The observation that a 20% Ni doped PMO, preactivated in H₂ stream provides higher activity and selectivity toward C7+ oxygenated than simple Ni20PMO, suggests that novel catalyst systems could be easily prepared and examined with careful variation of the Ni, Mg, and Al ratios and the addition of other dopants can be investigated to further tailor catalyst performance. In Ni20PMO, the formation of Ni nanoparticles upon activation in H₂ flow was confirmed by TEM analysis and EDX elemental mapping. Last, a simple hydrodeoxygenation protocol was established that enables the conversion of these mixtures into obtain fuel range alkanes with excellent carbon utilization (up to 89%).

Table 1. Weight and Carbon Yield of Alkanes Obtained from the HDO of Catalytic Coupling Mixtures from Acetone and 1-Butanol

Reaction Time (h)	Alkane Weight (mg)										Carbon Yield (%)	
	C3	C4	C5	C7	C8	C9	C10	C11	C12+	Total Alkanes	Total	C7+
2.5	2.3	0.5	4.7	71.5	0	2.2	6.2	212.2	34.2	333.8	78.4	60.0
5	11.9	0	9.7	73.6	2.7	0	7.7	223.9	37.2	366.5	86.0	63.9
10	8.3	0	9.0	68.7	0.4	2.6	7.1	231.1	39.9	369.2	86.6	66.2
20	3.4	0.7	15.2	60.0	0.8	2.4	7.5	244.7	44.1	378.7	88.9	70.5

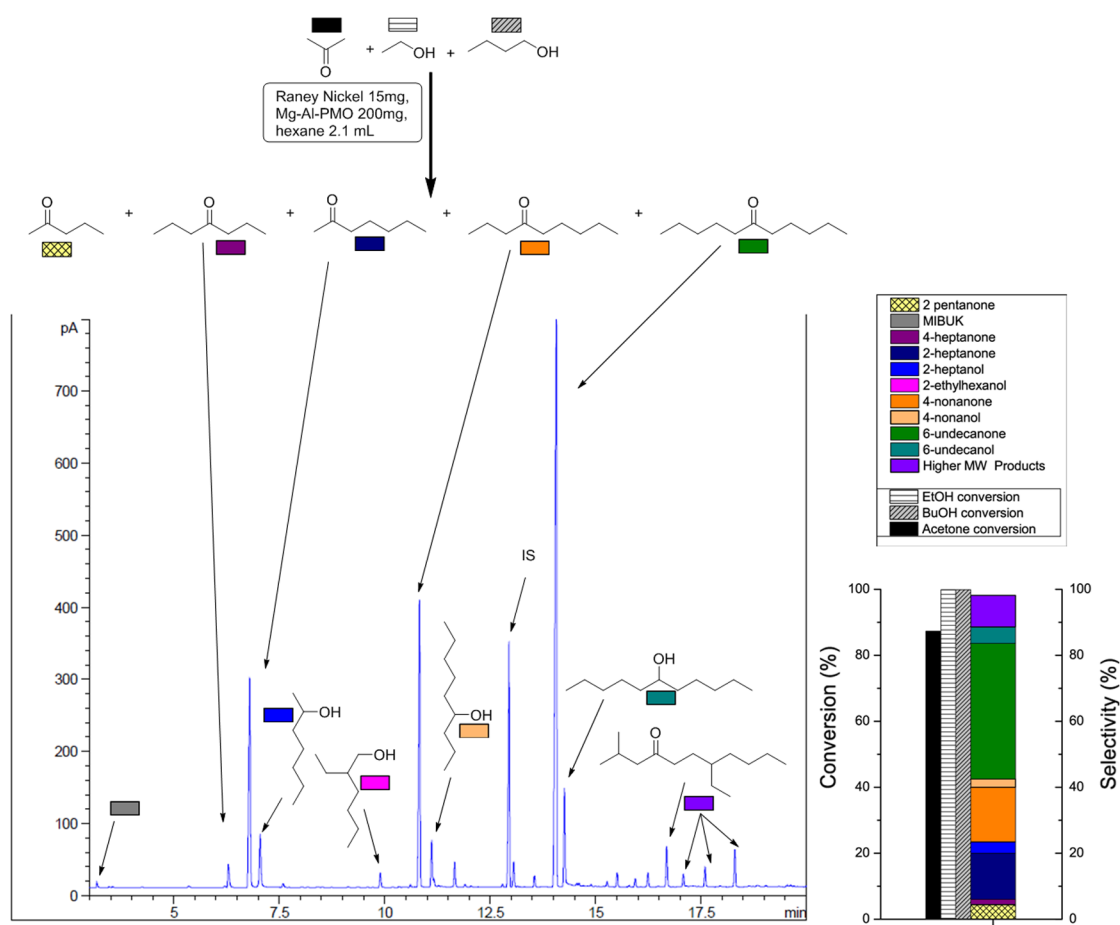


Figure 6. Results of the catalytic coupling of ABE mixture. Left: GC-FID trace of the product mixture showing all main components. Right: Product selectivities. Reaction conditions: 100 mg cat: Ni@Al₂O₃-SiO₂, dodecane (0.22 mmol, 50 μL, internal standard), 220 °C. Values determined by FID, using internal standard. For all data in table format, see the Supporting Information, Table S8.

Table 2. Weight and Carbon Yield of Alkanes Obtained from the HDO of Catalytic Coupling Mixtures from Acetone, Ethanol, and 1-Butanol

Reaction Time (h)	Alkane Weight (mg)										Carbon Yield (%)	
	C3	C4	C5	C7	C8	C9	C10	C11	C12+	Total Alkanes	Total	C7+
20	3.4	0.6	22.9	68.5	3.4	62.7	6.3	152.7	30.5	351.0	84.0	61.3

■ ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acssuschemeng.8b00733.

Additional details regarding the materials, experimental procedures, characterization of catalysts, analysis of products and calculations (PDF)

■ AUTHOR INFORMATION

Corresponding Author

*K. Barta. Email: k.barta@rug.nl.

ORCID

Marc C. A. Stuart: 0000-0003-0667-6338

Katalin Barta: 0000-0002-8046-4248

Notes

The authors declare no competing financial interest.

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